

1.

(

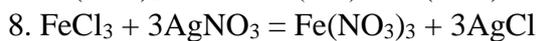
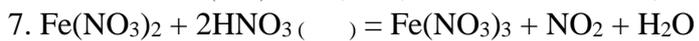
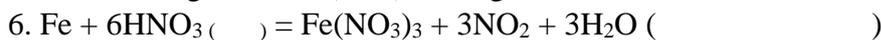
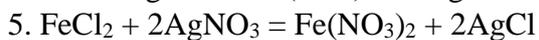
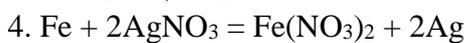
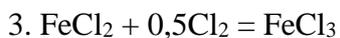
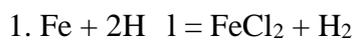
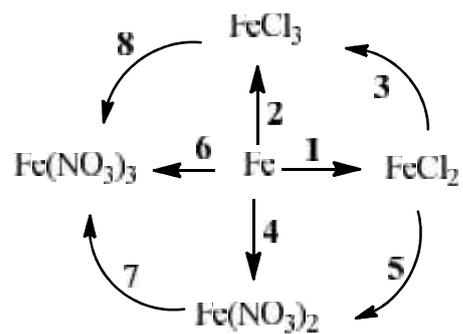
).

,

.

(20

)



2,5

2. 100 0.1 300 ° 338 (20)

$$PCl_5: n=0.1 / (5 \times 35.5 + 31) / = 4.8 \cdot 10^{-4} \quad (2)$$

$$PCl_5: PV=nRT, P=nRT/V = 4.8 \cdot 10^{-4} \times 62400 \times (300+273.15)/100 = 171.5 \quad (2)$$

$$PCl_5 = PCl_3 + Cl_2 \quad (3)$$

$$2 \times 171.5 = 343 \quad (2)$$

$$: 338 = (PCl_5) + PCl_3 \quad Cl_2$$

$$(PCl_5) + (PCl_3) = 171.5 \quad (3)$$

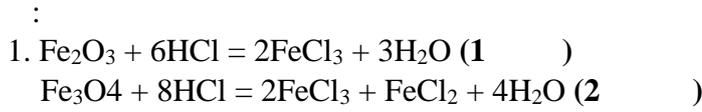
$$(Cl_2) = 338 - 171.5 = 166.5 \quad (PCl_3) = (Cl_2) = 166.5 \quad (3)$$

$$(PCl_5) = 171.5 - (PCl_3) = 5 \quad (3)$$

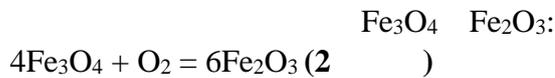
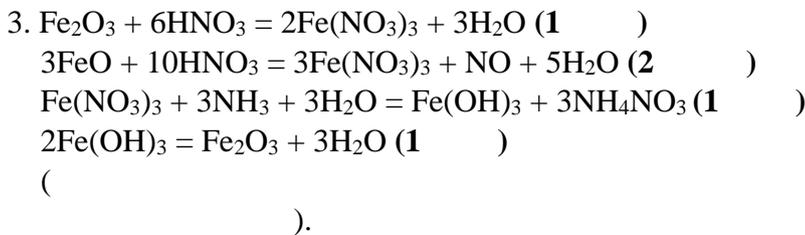
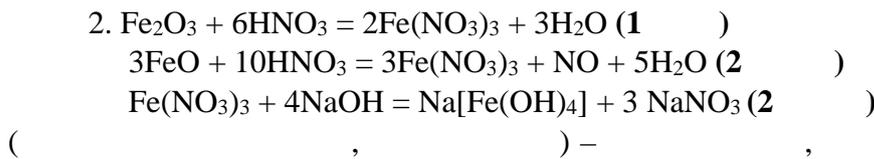
: PCl₅, PCl₃, Cl₂; 5, 166.5

3.

) : 1.50 2.04, 2.03, 2.04 ()
) . 1.50
 , 400 . 1.44, 1.43, 1.42 '
) 1.50
 900' . 1.53, 1.54, 1.54 '
 (20)



FeCl_2 Fe_2O_3 (c (III)) -

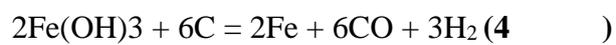
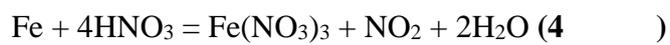
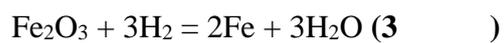
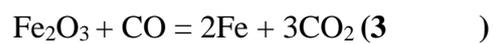


$4 \cdot 0,03/32 = 0,00375$ $- 4 \cdot 0,04/32 = 0,005$ (1)
 $- 0,00375 \cdot 232 = 0,87$ $- 0,005 \cdot 232 = 1,16$ (2)

71±13% (2)

4.

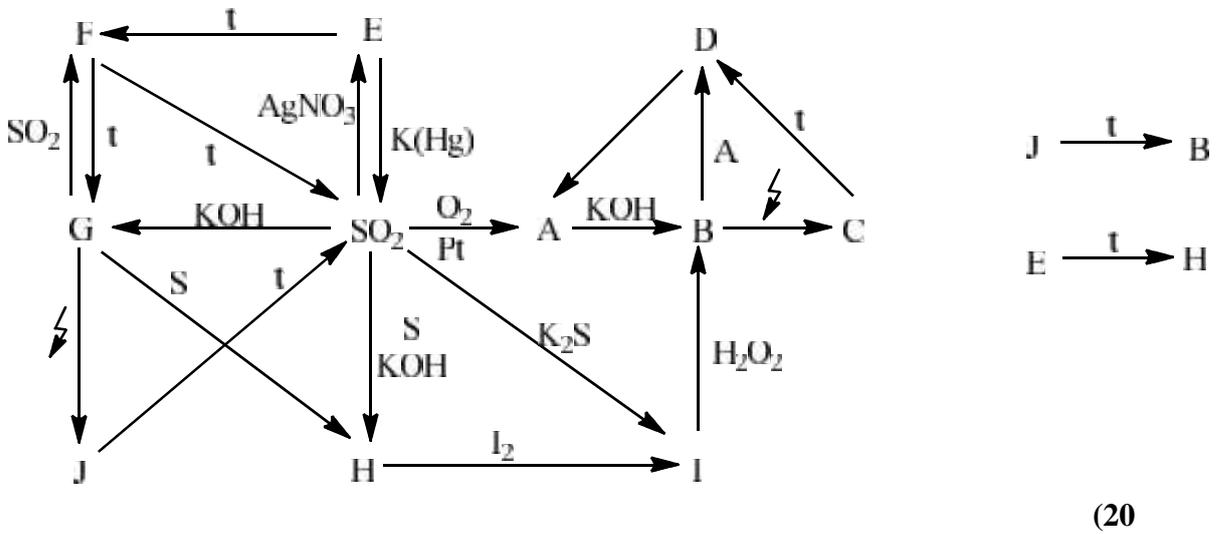
(20)



5.

A. B
 D. B
 C. B D
 E. E
 E G, F H. SO₂ F G
 SO₂ H. H
 J, SO₂ B I H G
 B () H. I
 -J.

| | C | D | E | F | H | I | J |
|--------|-------|-------|-------|-------|-------|-------|-------|
| (S), % | 23.70 | 25.20 | 31.07 | 28.83 | 33.68 | 42.38 | 26.89 |



C 23,70%.

K_xSO_y ,

$$M(C) = \frac{M(S)}{\omega(S)} = \frac{32}{0,237} = 135 \text{ г/моль}$$

$$135 = 39x + 16y + 32$$

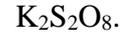
$$x=1 \quad y=4,$$



$$M(C) = \frac{2M(S)}{\omega(S)} = \frac{64}{0,237} = 270 \text{ г/моль}$$

$$270 = 39x + 16y + 64$$

$$x=2 \quad y=8,$$

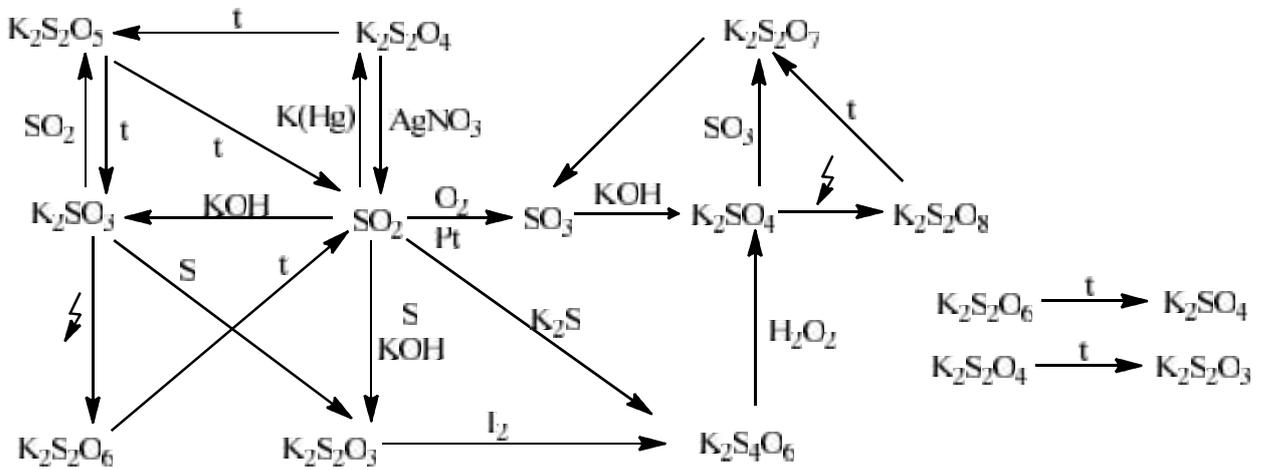


0,5

5

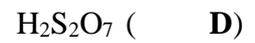
0,5

11



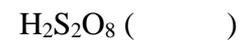
-J,

1.



(F)

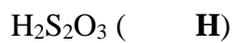
2.

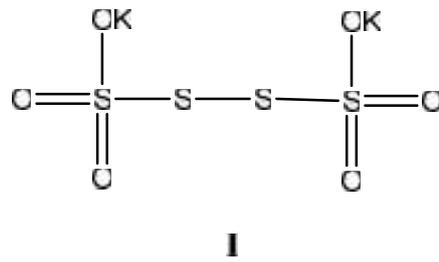
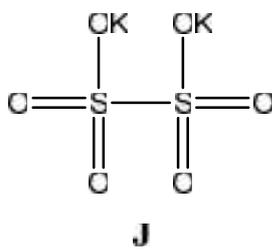
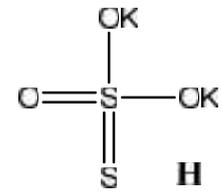
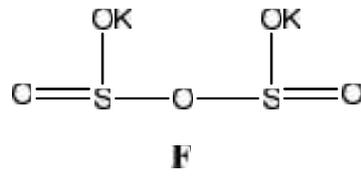
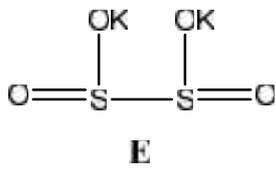
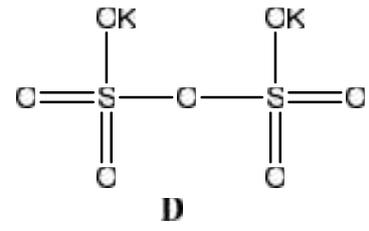
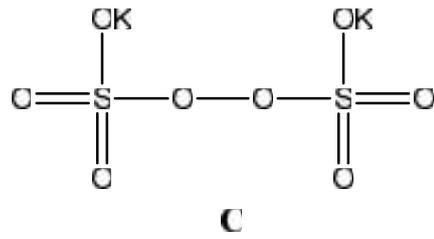
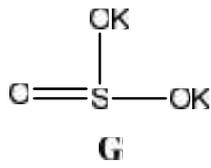


3.

(J, I)

4.





0,5
)

(

4